

Protocol



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Biomimetic tail-to-head terpene cyclizations using the resorcin[4]arene capsule catalyst

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Supplementary Information

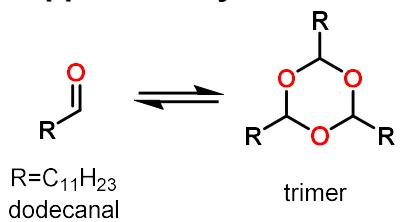


Fig S1 | Trimerization of dodecanal

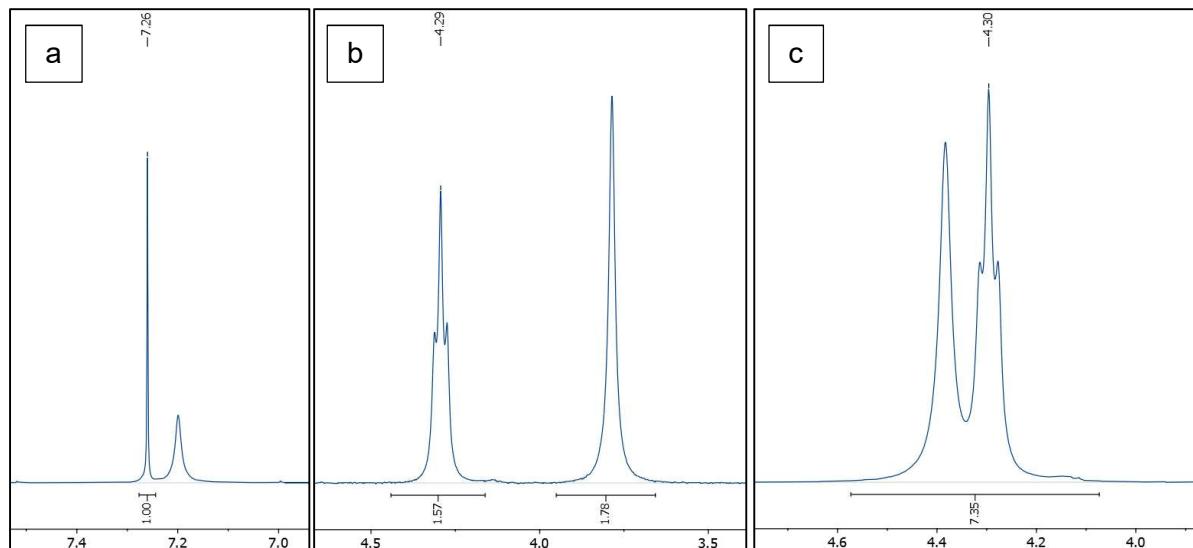


Fig. S2 | Example of ^1H NMR spectra for the determination of water in resorcin[4]arene (a) the CHCl_3 residual solvent is integrated first, (b) then the CH triplet (4.30 ppm) and the water signal (variable chemical shift). (c) If the CH triplet and the water singlet overlap, the aromatic CH singlet (6.11 ppm) should be used to deduce the water integral.

Derivation of the equations

Equation 2

$$c_M(\text{Et}_3\text{N titrant}) = \frac{n(\text{Et}_3\text{N})}{V(\text{volumetric flask})} = \frac{m(\text{Et}_3\text{N})}{V(\text{volumetric flask}) \times M_w(\text{Et}_3\text{N})} =$$

n (mol): molar amount

m (g): mass

V (L): volume

$M_w \left(\frac{\text{g}}{\text{mol}} \right)$: molar weight

$c_M \left(\frac{\text{mol}}{\text{L}} \right)$: molar concentration

Equation 3

$$c_M(HCl \text{ stock solution}) = \frac{n(HCl \text{ stock solution})}{V(HCl \text{ stock solution})} = \frac{n(Et_3N \text{ titrant})}{V(HCl \text{ stock solution})} = \frac{\frac{c_M(Et_3N \text{ titrant}) \times V(Et_3N \text{ titrant})}{V(HCl \text{ stock solution})}}{V(HCl \text{ stock solution})}$$

V (μL): volume

n (mol): molar amount

$c_M \left(\frac{\text{mol}}{\text{L}} \right)$: molar concentration

Equation 4

$$\begin{aligned} r \left(\frac{H_2O \text{ present}}{\text{capsule}} \right) &= \frac{n(H_2O \text{ in sample}) - n(H_2O \text{ in } CDCl_3)}{n(\text{resorcin[4]arene in sample})} \times 6 \\ &= \frac{\frac{\text{int}(H_2O \text{ in sample})}{2} - \frac{\text{int}(H_2O \text{ in } CDCl_3)}{2}}{\frac{\text{int}(CHCl_3 \text{ in sample})}{4} - \frac{\text{int}(CHCl_3 \text{ in } CDCl_3)}{4}} \times 6 \end{aligned}$$

r : ratio

n : amount

int : peak integral

Equation 5 and 10

$$\begin{aligned} m(H_2O \text{ to add}) &= n(H_2O \text{ to add}) \times Mw(H_2O) = \frac{r \left(\frac{H_2O \text{ to add}}{\text{capsule}} \right) \times n(RS) \times Mw(H_2O)}{6} = \\ &= \frac{r \left(\frac{H_2O \text{ to add}}{\text{capsule}} \right) \times m(RS) \times Mw(H_2O)}{Mw(RS) \times 6} \\ &= \frac{\left(r \left(\frac{H_2O \text{ target}}{\text{capsule}} \right) - r \left(\frac{H_2O \text{ present}}{\text{capsule}} \right) \right) \times m(RS) \times Mw(H_2O)}{Mw(RS) \times 6} \end{aligned}$$

n (mol): molar amount

m (g): mass

r (-): ratio

$Mw \left(\frac{\text{g}}{\text{mol}} \right)$: molar weight

RS : resorcin[4]arene

Equation 6

$$r \left(\frac{H_2O}{capsule} \right) = \frac{n(H_2O \text{ in sample})}{n(resorcin[4]arene)} \times 6 = \frac{\frac{int(H_2O \text{ in sample})}{2}}{\frac{int(CH - resorcin[4]arene)}{4}} \times 6$$

r: ratio

n (mol): molar amount

int: peak integral

Equation 7

$$\begin{aligned} r \left(\frac{H_2O}{capsule} \right) &= \frac{n(H_2O) \times 6}{n(RS)} = \frac{m(H_2O) \times 6}{n(RS) \times M_w(H_2O)} = \frac{ppm(H_2O) \times m(sol) \times 6}{10^6 \times n(RS) \times M_w(H_2O)} \\ &= \frac{ppm(H_2O) \times m(sol) \times M_w(RS) \times 6}{10^6 \times m(RS) \times M_w(H_2O)} \\ &= \frac{ppm(H_2O) \times m(sol) \times Mw(RS) \times 6}{10^6 \times V(sol) \times M_w(H_2O) \times c_w(RS)} \end{aligned}$$

r (-): ratio

m (g): mass

n (mol): molar amount

V (mL): volume

Mw ($\frac{g}{mol}$): molar weight

c_w ($\frac{g}{mL}$): weight concentration

RS : resorcinarene

sol : solution

NMR spectra

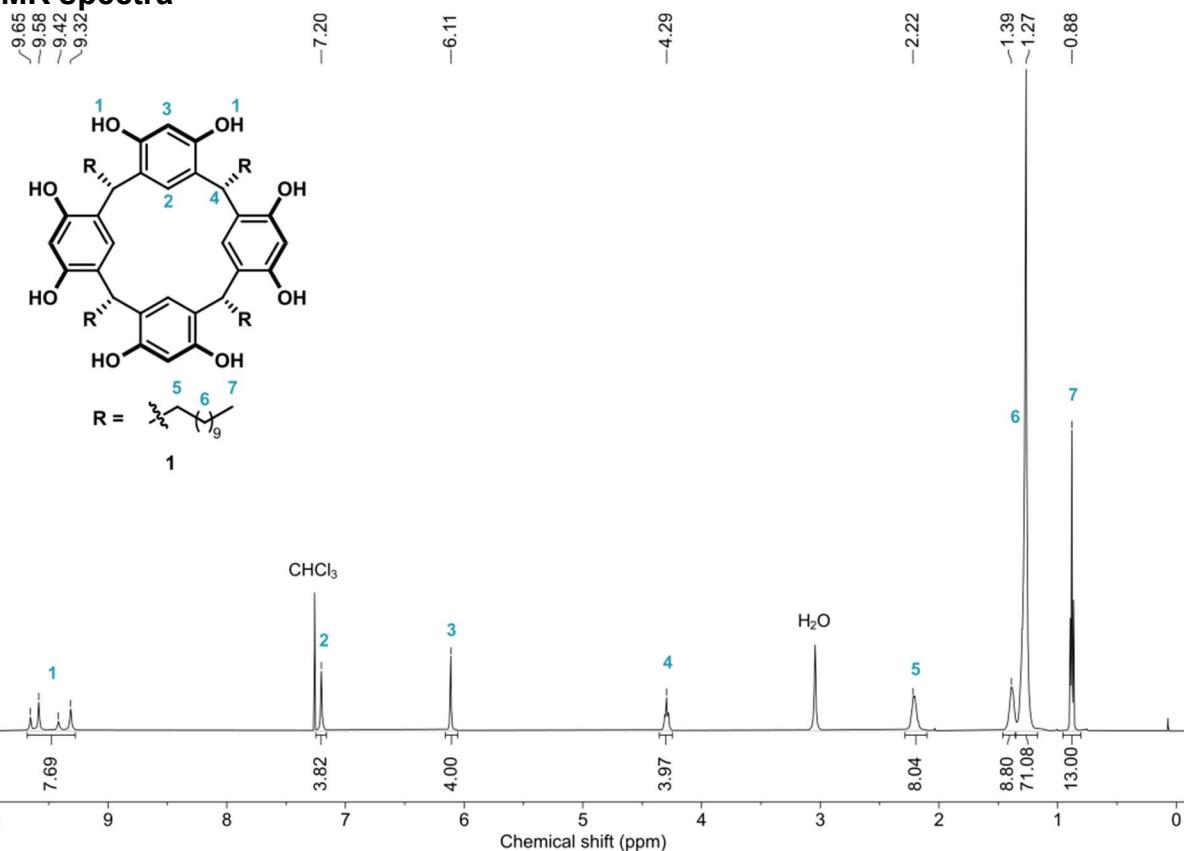


Fig. S3 | ¹H NMR spectrum of resorcin[4]arene **1** (500 MHz, chloroform-d, 298K)

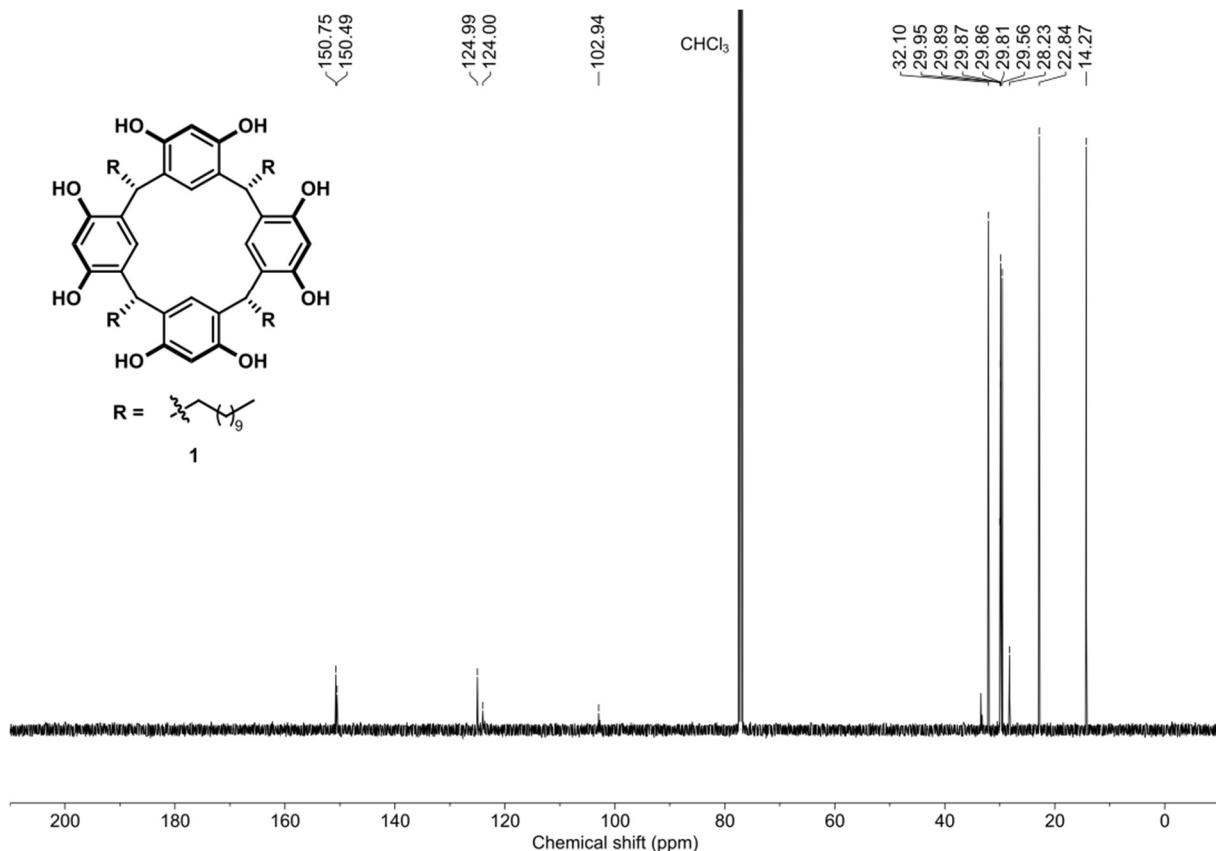


Fig. S4 | ¹³C NMR spectrum of resorcin[4]arene **1** (126 MHz, chloroform-d, 298K)

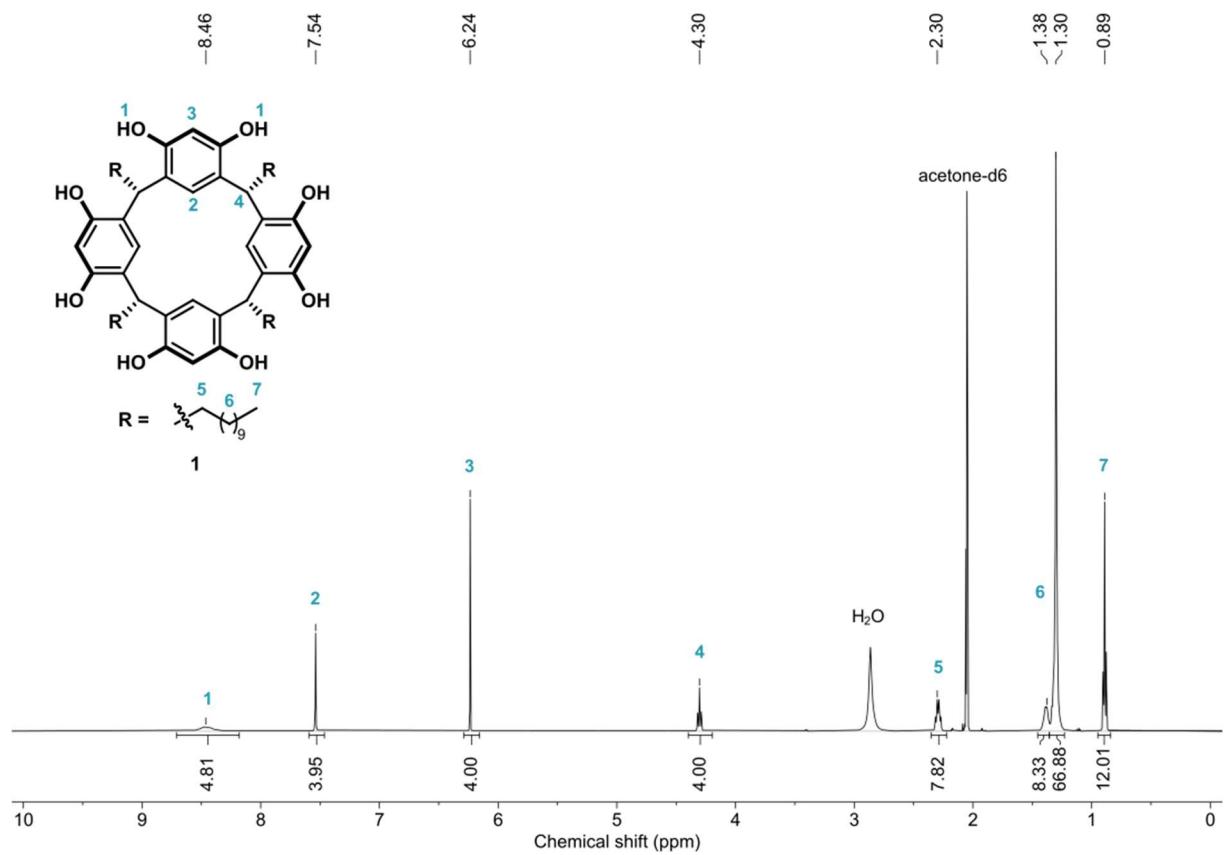


Fig. S5 | ^1H NMR spectrum of resorcin[4]arene **1 (500 MHz, acetone-d₆, 298K)**

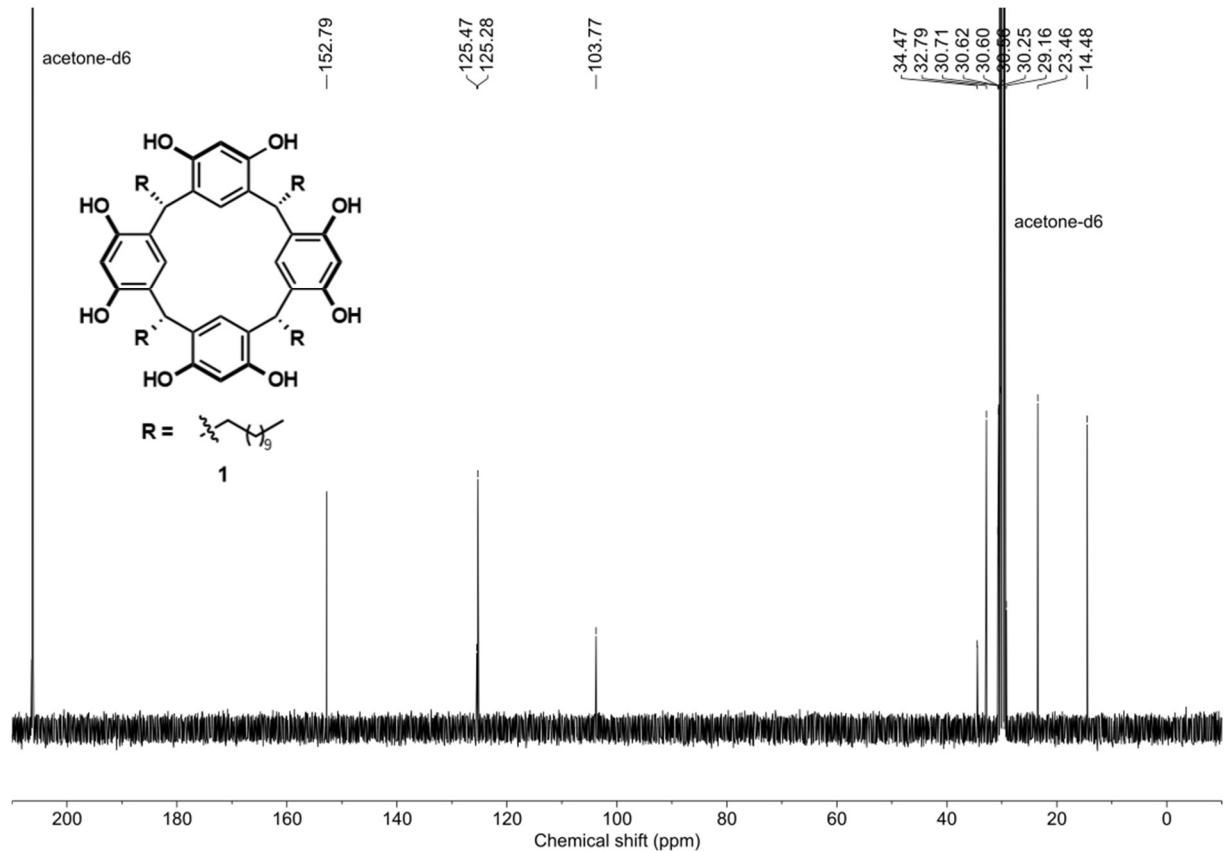


Fig. S6 | ^{13}C NMR spectrum of resorcin[4]arene **1** (126 MHz, acetone-d₆, 298K)

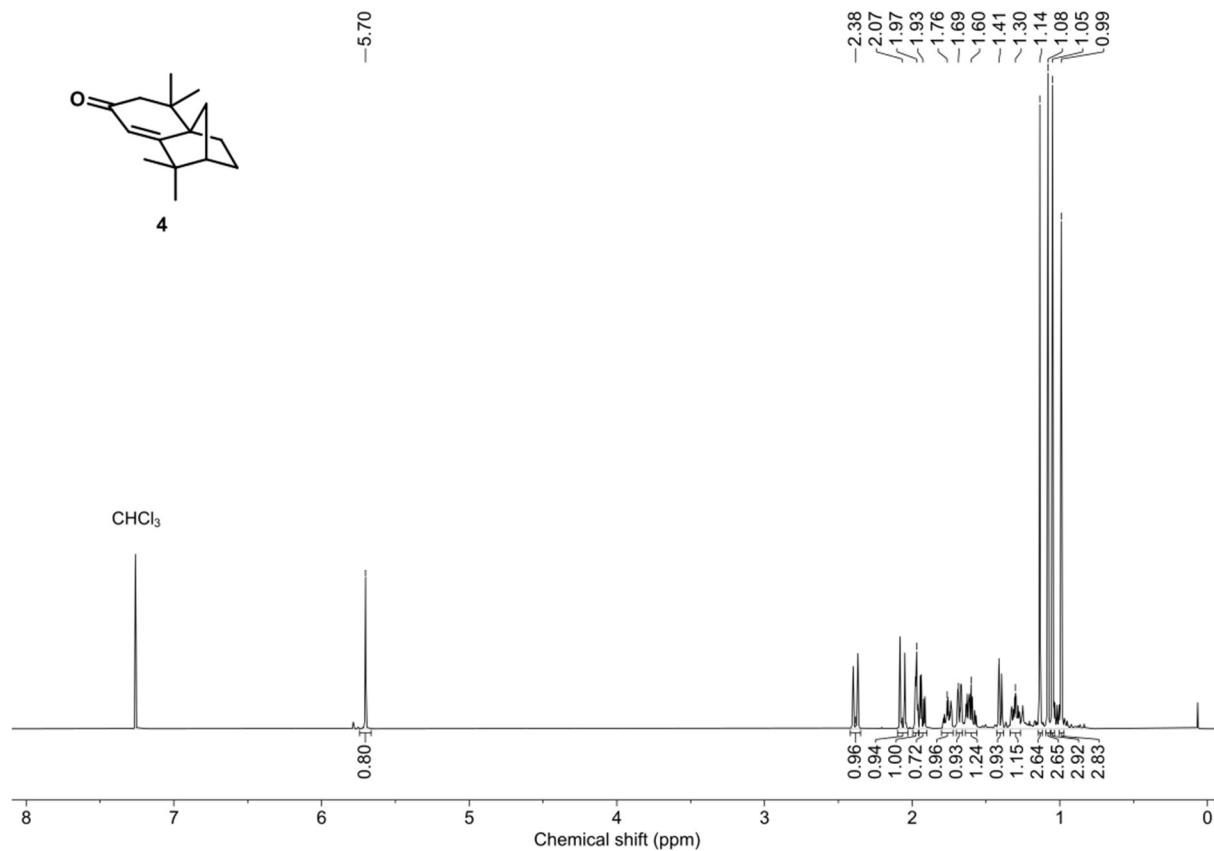


Fig. S7 | ^1H NMR spectrum of isolongifolenone **4** (500 MHz, chloroform-d, 298K)

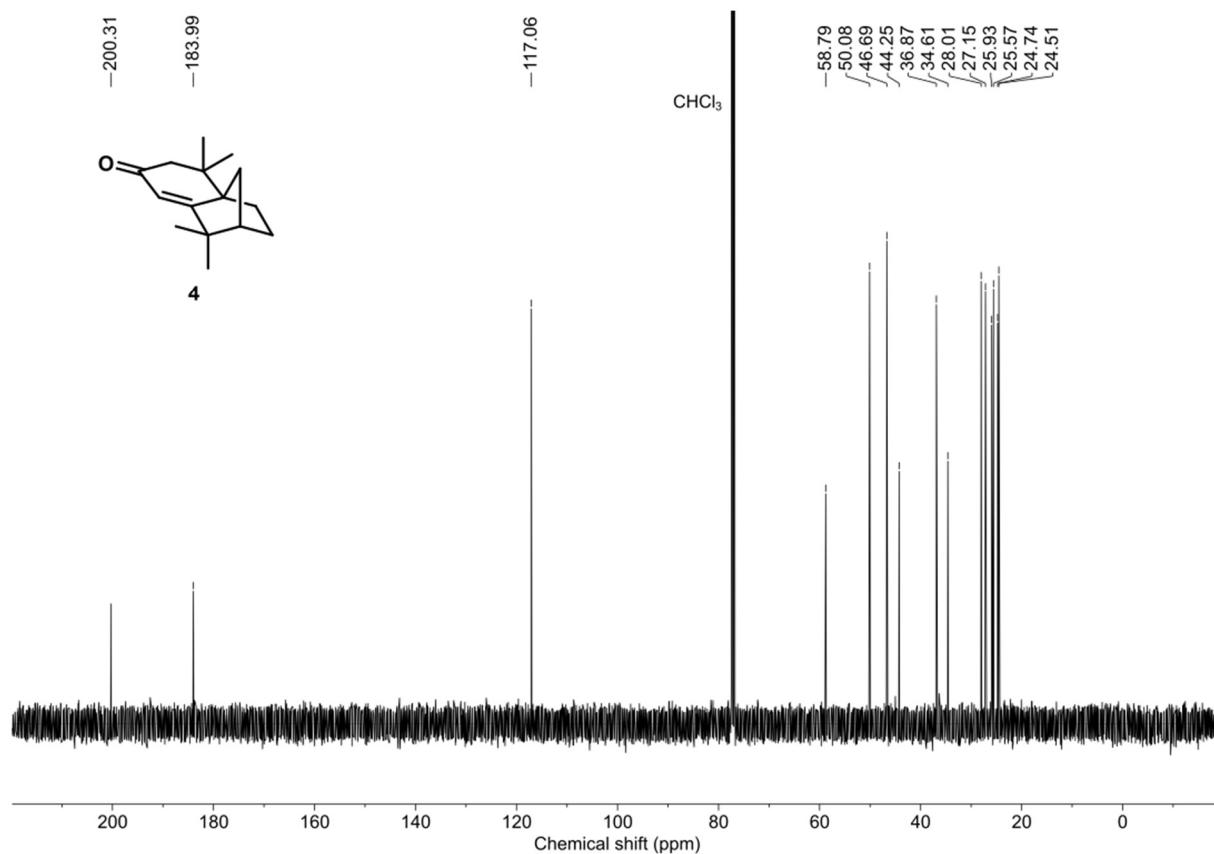


Fig. S8 | ^{13}C NMR spectrum of isolongifolenone **4** (126 MHz, chloroform-d, 298K)

